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Edit. ... Atoms of the same element with

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different numbers of neutrons are called \_\_\_\_\_. answer choices . radioactive elements. metalloids. metals.

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Chapter 1. Atoms Student Objectives 1.1  
A Particulate View of the World:  
Structure Determines Properties •  
Define atoms, molecules, and the

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science of chemistry. • Represent a simple molecule, water, using spheres as atoms. 1.2 Classifying Matter: A Particulate View

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Electrons are considered valence electrons if they are in the outermost

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electron shell. Electrons removed from atoms always increase the stability of the atom. Electrons circle the nucleus in...

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The proton is positive, the electron is negative, and the neutron is electrically

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neutral. Protons and neutrons have about the same mass. Each is about 2,000 times the mass of an electron. Since protons and neutrons exist in the nucleus, almost all the mass of an atom is concentrated there.

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## **Chapter Test Atoms And The Periodic Table Answers**

d) an atom or group of atoms with a net positive charge. 16. A cation is defined as a) a charged atom or group of atoms with a net negative charge. b) a stable

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atom. c) a group of stable atoms. d) an atom or group of atoms with a net positive charge. 17. Atoms of the same element with different mass numbers (or number of neutrons) are called a ...

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Chapter 5 TEST: The Periodic Table  
Multiple Choice Identify the choice that

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best completes the statement or answers the question. \_\_\_\_ 1. The order of elements in the periodic table is based on a. the number of protons in the nucleus. b. the electric charge of the nucleus. c. the number of neutrons in the nucleus. d. atomic mass. \_\_\_\_ 2.

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Chapter 1. Electronic Structure and Chemical Bonding ... Many of the physical and chemical properties of elements can be correlated to their unique electron configurations. The valence electrons, electrons in the outermost shell, are the determining factor for the unique chemistry of the element. ... which determine the

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chemical properties of ...

## **1.1: Electronic Configuration of Atoms - Chemistry LibreTexts**

38 Properties of Atoms and the Periodic  
Table Chapter Test B (continued) 19.

According to present atomic theory, the  
location of an electron in an atom is best  
represented by a probability cloud. 20.

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As you move from left to right in a row of the periodic table, metallic properties increase.

**001 052 C16 CR 896260 3/24/10**

**4:44 AM Page 37 U-S083 113 ...**

Chapter Review Date Properties of  
Atoms and the Periodic Table Class Part  
A. Vocabulary Review Directions: On the

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space at the left, write the term that correctly completes each statement. Use each term once. metals groups electrons chemical symbol isotopes metalloids nucleus average atomic mass transition elements mass number electron cloud

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- Atoms of an element are similar to each other but different from those of other elements.
- All atoms of a particular element contain the same numbers of electrons and protons.
- Chemical compounds are made up of specific combinations of atoms in specific ratios.

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**Chapter 2—Atoms and the Periodic Table**

Introduction; 18.1 Periodicity; 18.2 Occurrence and Preparation of the Representative Metals; 18.3 Structure and General Properties of the Metalloids; 18.4 Structure and General Properties of the Nonmetals; 18.5 Occurrence, Preparation, and Compounds of

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Hydrogen; 18.6 Occurrence, Preparation, and Properties of Carbonates; 18.7 Occurrence, Preparation, and Properties of Nitrogen

## **Answer Key Chapter 10 - Chemistry: Atoms First 2e | OpenStax**

Use the concepts in this chapter to obtain an estimate for the number of

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atoms in the universe. Make the following assumptions: (a) All of the atoms in the universe are hydrogen atoms in stars. (This is not a ridiculous assumption because over threefourths of the atoms in the universe are in fact hydrogen.

## **Atoms | Chemistry Structure and**

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compounds is a review of theories presented by Dalton, Thomson, Rutherford, and Millikan.

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Most molecules are made of atoms of different elements, such as water. But a molecule may also be made of atoms of

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the same element, such as those shown in Figure 5. A compound is made of atoms of two or more different elements, but a molecule may be of the same elements or different elements.

molecules 40 CHAPTER 2 Figure 4 A water molecule ...

## **CHAPTER 2 Matter**

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PDF Assessment Chapter Test A -  
clarkchargers.org 22. Answers should  
include three of the following: All matter  
is composed of extremely small particles  
called atoms. All atoms of an element  
are identical in size, mass, and other  
properties; atoms of different elements  
differ in size, mass and other properties.

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